

Removal of iron and manganese from groundwater: a study of using potassium permanganate and sedimentation

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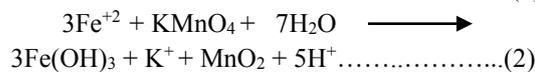
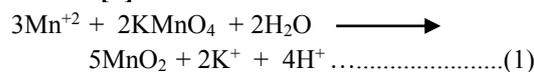
Abstract. Experiments are done for different combinations of Fe⁺² and Mn⁺² concentrations. The obtained results show that Potassium Permanganate (PP) gives good results in iron and manganese removal. By using PP dose near to half of the theoretically required one, it can remove up to 100 % and 90 % of iron and manganese, respectively over different tested concentrations at pH=7.0. Increasing rate of filtration influences the Mn⁺² removal process obviously. Sedimentation is required when combined concentrations of iron and manganese are greater than 5.0 ppm to reduce filter rapid clogging. Using conventional treatment with adding alum, flocculation, sedimentation and filtration can remove up to 97% and 18% of iron and manganese, respectively. Using PP in addition to alum enhances manganese removal but decreases iron removal. However, using alum with increasing pH to 10 leads to 100 % and 95 % of Fe⁺² and Mn⁺² removal and increases filter working period.

1. INTRODUCTION

Groundwater sources in Egypt can be divided into renewable aquifers like Nile valley and Delta and non-renewable aquifers like the aquifer of the western desert in the Nubian sandstone [1]. Iron and manganese are usually present in groundwater as dissolved minerals or associated with other components [2]. Existence of iron and manganese in water causes many problems like water coloring and taste, clothes staining and encouraging bacterial growth in water distribution networks which affect the pipes transfer efficiency [3] but in general they do not cause health problems [4]. The secondary maximum contaminant levels for iron and manganese are 0.3 mg/l and 0.05 mg/l, respectively [5]. The most common methods used to remove iron and manganese include oxidation by aeration, chlorine, chlorine dioxide, potassium permanganate and/or ozone followed by filtration alone or sedimentation and filtration [6,7]. There are other methods that can also be used like using filters with special media like

green sand, using ion exchange, biological methods or membrane filtration [8].

Potassium Permanganate (PP) is considered a stronger oxidant and has many advantages over other oxidants. Oxidation chemistry of iron and manganese by PP can be described as follows [9]:



To oxidize one mg of iron and one mg of manganese, about 0.94 mg and 1.92 mg of PP are required, respectively [10].

When iron and manganese exist in high concentrations, the filter runs less than 24 hours so a clarification step is needed before filtration to increase the filtration period [11]. So, treatment method would there include flocculation, sedimentation and filtration stages. [12]

This paper studies using PP followed by filtration only or using PP and/or alum (Al) followed by flocculation, sedimentation and

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filtration for Fe^{+2} and Mn^{+2} oxidation and removal.

2. MATERIALS AND METHODS:

2.1. Study Method:

Simulated groundwater was prepared by adding salts of iron and manganese to tap water. The study consists of three groups of experiments: the first one discusses the different factors (e.g. dosages, detention time, pH) that affect the oxidation of Fe^{+2} and Mn^{+2} by using PP followed by direct filtration. Experiments are done for Fe^{+2} and Mn^{+2} = 1.50 and 1.0 mg/l, respectively which represent iron and manganese concentrations in the Delta region, Egypt. The second group investigates the results of the first one on other combined concentrations of Fe^{+2} and Mn^{+2} by using direct filtration alone. The third group discusses the efficiency of using sedimentation if high concentrations of Fe^{+2} and Mn^{+2} exist. This phase includes flocculation, sedimentation and Filtration.

2.2. Pilot plant:

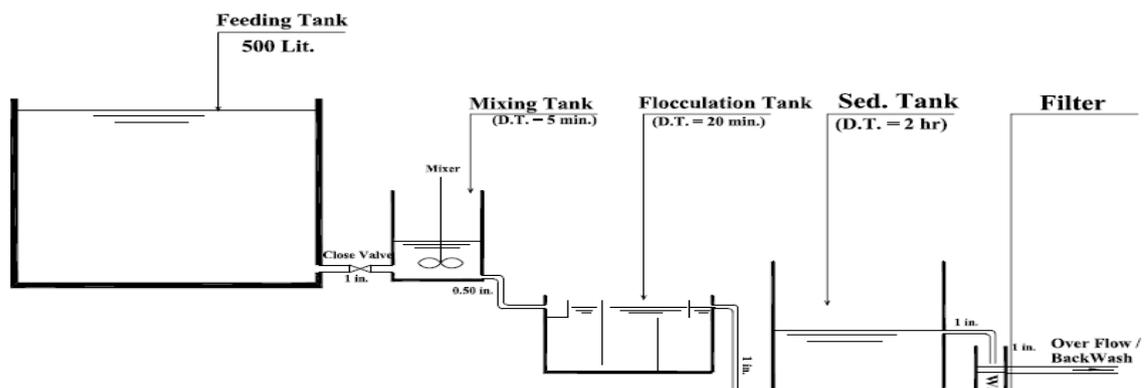


Fig. (1): Pilot Plant with mixing, flocculation, sedimentation and filtration tanks. When direct filtration mode is used, the flocculation and sedimentation tanks are removed.

Figure (1) shows the pilot plant, which is constructed for the study. It consists of feeding tank, process tank for adding and mixing chemicals, flocculation and sedimentation tanks and a Rapid Sand Filter (RSF). For 1st and 2nd phase experiments, the flocculation and sedimentation tanks are not

included. The RSF is made of PVC pipe of internal diameter = 100 mm and includes 35 cm gravel layer with diameter 6:25 mm and 75 cm of coarse sand with diameter 1.18:1.60 mm. The Rate of Filtration is obtained by controlling the outlet filter valve.

2.3. Chemicals:

The simulated groundwater containing iron and manganese is made by adding ferrous sulfate heptahydrate ($FeSO_4 \cdot 7H_2O$) and manganese sulfate mono-hydrate ($MnSO_4 \cdot H_2O$) to tap water. Potassium permanganate ($KMnO_4$) with 99.90% purity and aluminum sulphate hexadecahydrate ($Al_2(SO_4)_3 \cdot 16H_2O$) were used as sources of PP and alum. Sodium hydroxide ($NaOH$) was used to adjust pH.

2.4. Devices and Analyses:

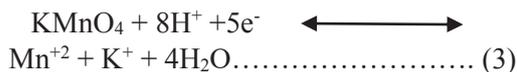
Iron and manganese measure devices were used to measure iron and manganese concentrations (Hanna, USA). Portable pH device was used to measure water pH. The devices were calibrated before the study.

3. RESULTS AND DISCUSSIONS:

3.1. Group 1: study using Potassium Permanganate:

In this group, the experiments were done for Fe^{+2} and Mn^{+2} concentrations = 1.50 and 1.0 mg/l, respectively at constant ROF= 150 $m^3/m^2/d$, discuss the removal efficiency of iron and manganese using potassium permanganate under different conditions like different dosages of PP, contact time, effect of pH and ROF effect. The experiments lasted for 6 hours before taking samples.

Effect of PP dosages: Fig. (2) shows the results of using different PP dosages for Fe^{+2} and Mn^{+2} removal. Using a dosage of 1 ppm of PP can remove up to 97% of iron just after 10 minutes. For manganese, using PP enhances the removal process greatly at pH=7.0. Using PP dose = 2.0 ppm -which is near to half of the theoretically calculated dose- can remove 66 % of manganese after 10 minutes. Using doses near to the theoretical dose remove up to 80% of manganese in just 10 minutes. Increasing PP dosage more than the theoretical one has a bad effect on water. When the dosage of 4 ppm is used, the water is colored pink from the effect of increased PP dosage. The increased PP quantity contains manganese according to the following equation [13]:



Therefore, a special care should be taken into consideration for choosing the PP dosage. These results agree with other studies that recommend using PP dosages near the theoretical ones [13].

Effect of Retention Time: The results for using PP dose =2.0 ppm for RT=5 to 20 minutes are shown in Fig. (3). The oxidation process of iron using PP happens very fast. Complete iron oxidation needs less than 5 minutes. The oxidation process of manganese using PP also happens fast and is enhanced by increasing RT. At RT less than 5 minutes, about 66% of manganese is oxidized, however, at RT equals to 20 minutes increases the RR to 90%.

Effect of pH: Fig. (4) shows results of pH effect on Fe^{+2} and Mn^{+2} oxidation by PP dose = 2.0 ppm and RT = 10 min. It is shown that iron is oxidized at pH greater than 7.0 and much of manganese oxidation by using PP happens also at pH near to 7.0. Increasing pH to 8 and 9, increases the RR to 75% and 85%, respectively. Therefore, it is concluded that pH affects the process of Fe^{+2} and Mn^{+2} oxidation by PP slightly or in other words, they need pH greater than 7.0.

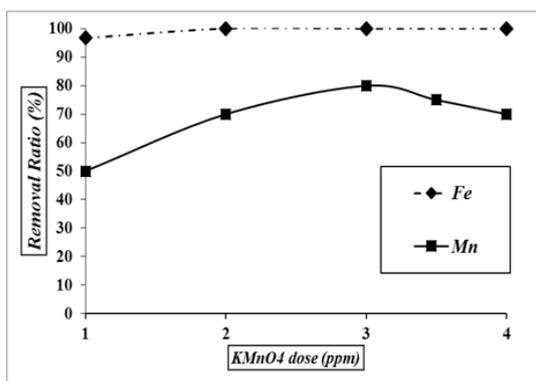


Fig. (2): Effect of using different PP doses on Fe^{+2} and Mn^{+2} removal (Initial Fe^{+2} and Mn^{+2} are 1.50 and 1.0 mg/l, pH=7.0, ROF=150 m/d and RT=10 min.)

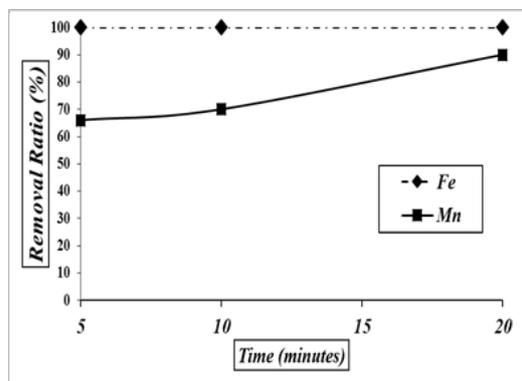


Fig. (3): Effect of retention time on Fe^{+2} and Mn^{+2} removal by PP (Initial Fe^{+2} and Mn^{+2} are 1.50 and 1.0 mg/l, pH=7.0, PP=2.0 ppm and ROF=150 m/d)

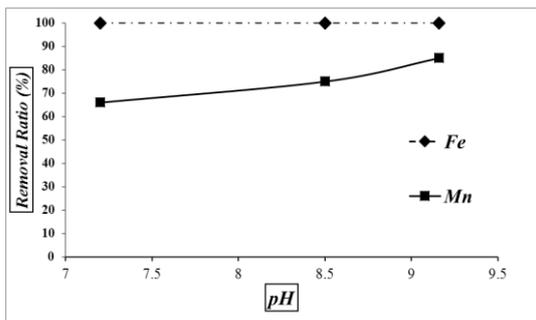


Fig. (4): Effect of pH on Fe^{+2} and Mn^{+2} removal by PP (Initial Fe^{+2} and Mn^{+2} are 1.50 and 1.0 mg/l, PP=2.0 ppm ROF=150 m/d and RT=10 min.)

The obtained results agree with other studies which found that the required dose of PP to oxidize Mn^{+2} was less than that indicated by stoichiometry. It is thought that when Mn^{+2} is separated on the filter, it starts to coat the filter sands and convert it to work like green sand filter so the required dose becomes smaller [10]. The oxidation time ranges from 5 to 10 minutes, provided that the pH is over 7.0 [12]. On the other hand, some studies found that the required dosage is slightly more than the required theoretical dose at pH less than 8.0 [13] and so, the required dose should be determined accurately.

Rate Of Filtration (ROF) effect: Fig. (5) shows the results of experiments that discuss the effect of increasing ROF on Fe^{+2} and Mn^{+2} removal. PP dose = 2.0 ppm was used for RT = 20 min. For iron: increasing ROF from 150 to 180 m/d does not affect the removal efficiency. At ROF=220 m/d, RR becomes

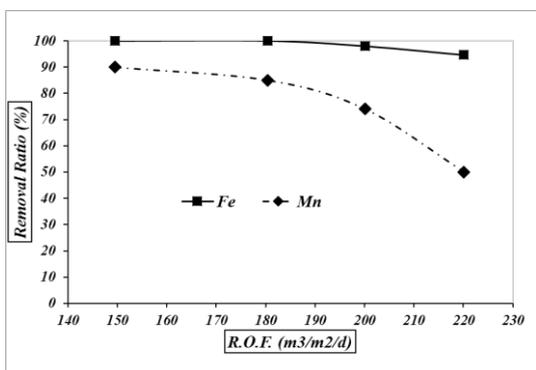


Fig. (5): Effect of R.O.F. on Fe^{+2}/Mn^{+2} removal (Initial Fe^{+2}/Mn^{+2} conc. is 1.50/1.0 mg/l, pH=7.0, P.P.=2.0 ppm and D.T.=20 min)

Table (1): Fe^{+2}/Mn^{+2} concentrations of phase 2 experiments

Fe^{+2} (mg/l)	Mn^{+2} (mg/l)	Combined Conc. (mg/l)	PP dose (ppm)
1.5	1	2.5	2
3	2	5	3.5
5	3	8	5.5
6	4	10	7

95%. But in general, it is seen that iron removal is slightly affected by increasing ROF. For manganese: the biggest RR is obtained at ROF equals 150 m/d. At ROF equals 220 m/d, the removal ratio becomes 50% only, therefore, it is obvious that manganese removal depends on ROF in contrast to iron. Therefore, ROF of 150-180 m/d (6.25-7.5 m/h) is recommended for Mn^{+2} and 150-220 m/d (6.25-9.17) for iron. However, it is reported that the recommended ROF for Mn^{+2} removal is about 15-18 m/hr where for iron is 6-7.5 which is considered totally different from our study [12].

3.2. Group 2: Study results of Phase 1 on other Fe^{+2} and Mn^{+2} concentrations:

In this group, experiments were done for different concentrations of iron and manganese with using half of the required stoichiometric dose of PP. Table (1) shows these concentrations and the used PP dose. The experiments started with ROF equals to 180 m³/m²/d, pH=7.0 and RT equals to 20 minutes. The experiments lasted for 24 hours.

Effect of using PP: Fig. (6) shows results of using half of the stoichiometric PP dose on the removal of Fe^{+2} and Mn^{+2} different concentrations. The figure shows that using PP is very efficient to oxidize and remove Fe^{+2} and Mn^{+2} at different concentrations, where using half of stoichiometric dose of PP can remove about 98% of Fe^{+2} and more than 90% of Mn^{+2} despite increasing combined concentrations to 10 mg/l.

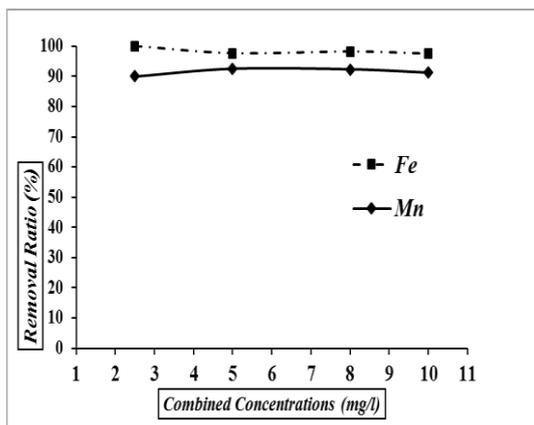


Fig. (6): Effect of using half dose of PP on removal of different concentrations of Fe^{+2} and Mn^{+2} (Initial pH=7.0, RT=20 min. Samples are taken and analyzed after 8 hours from filter run start).

Filter clogging rate: Fig. (7) shows the change in ROF with time (in hours) for experiments of group 2. The change in ROF expresses the velocity of filter clogging at different concentrations. The figure shows that the ROF decreases with time as expected. This occurs as removed iron and manganese accumulate on filter media and start to clog the filter. It is shown that increasing Fe^{+2} and Mn^{+2} concentrations lead to rapid filter clogging. When concentrations of iron and manganese exceed 5.0 mg/l, the removed particles cause rapid clogging of filters and decrease the period of filter run to be less than 12 hours. This period is not practical nor economical to be applied to water treatment plants as the ratio of backwash water will increase. And so, it is recommended that when these concentrations exceed 5.0 mg/l to use sedimentation before filtration. Other studies suggest using sedimentation when combined concentrations of Fe^{+2} and Mn^{+2} exceed 8.0 mg/l [8] or when iron concentration is greater than 5.0 mg/l [12].

3.3. Group 3: Study using conventional treatment to remove High concentrations of Fe^{+2} and Mn^{+2} :

This group of experiments deals with using conventional treatment to remove Fe^{+2} and Mn^{+2}

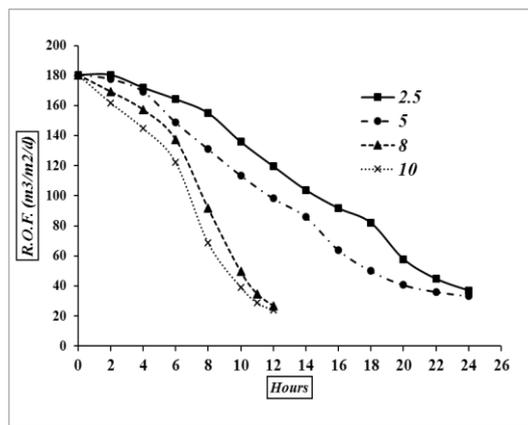


Fig. (7): Change in ROF of different concentrations of Fe^{+2} and Mn^{+2} (Initial pH=7.0, RT=20 min. All experiments Started at ROF=180 m/d)

when found in high concentrations. The treatment consists of adding alum (with/without other chemicals), mixing with water for 5 minutes, flocculation for 25 min., sedimentation for 2.0 hrs and then filtration. This set discusses using alum alone, alum with PP and alum with increasing pH.

Fig. (8) shows the results of these experiments. It is shown that using alum alone can remove high ratio of Fe^{+2} . By using alum dose=60 ppm, about 97% of Fe^{+2} can be removed. But, RR of Mn^{+2} does not exceed 18%. Using PP with alum enhances Mn^{+2} RR to reach about 63% but decreases Fe^{+2} RR to 79% when PP dose=5.0 ppm used with alum dose=40 ppm. using alum alone with increasing water pH to over 10 leads to complete removal of iron and about 95% of Mn^{+2} .

Another study found that Fe^{+2}/Mn^{+2} removal by coagulation and clarification were about 18:75% and 8:24% for Fe^{+2}/Mn^{+2} , respectively for different concentrations of both. However, when PP was used, Fe^{+2} RR was about 99% which is greater than our study results and Mn^{+2} RR was 72% which are like the above results [14].

4. CONCLUSIONS:

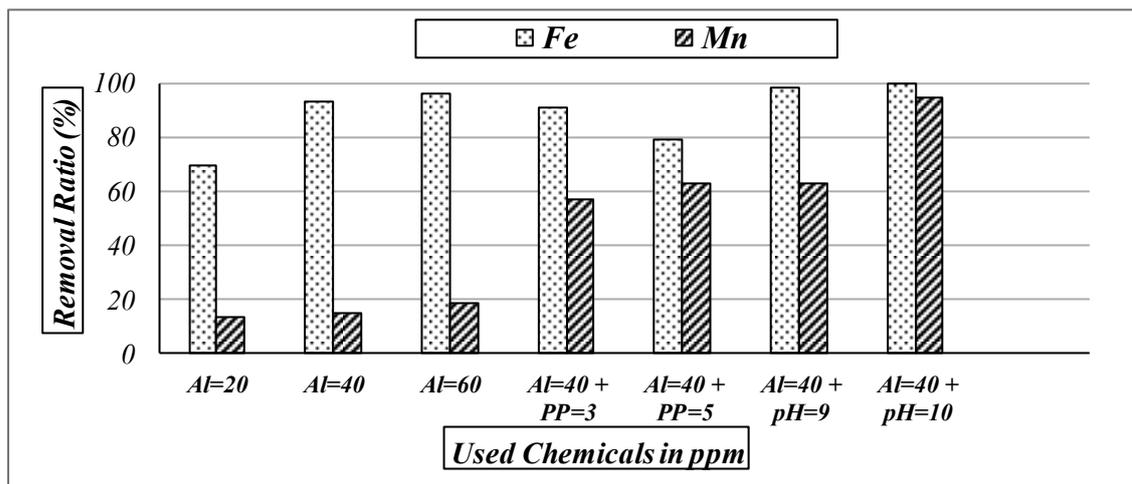


Fig. (8): Results of using conventional treatment for removal of $Fe^{+2}=5.0$ mg/l and $Mn^{+2}=3.0$ mg/l (Initial pH=7.0, All experiments Started at ROF=150 m/d and samples are taken after 10 hours from filtration start).

Iron and manganese can be removed from groundwater by using oxidation by potassium permanganate followed by filtration which gives good results. By using doses near to half of the theoretically calculated dose, it can remove up to 100% and 90% of iron and manganese at pH=7.0 over tested concentrations. Sedimentation is required when combined concentrations of iron and manganese are greater than 5.0 ppm to increase the filtration period. Using conventional treatment of adding alum, flocculation, sedimentation and filtration can remove up to 97% and only 18% of iron and manganese, respectively. Using PP in addition to alum enhances manganese removal but decreases iron removal. However, using alum with increasing pH to 10 leads to 100% and 95% of Fe^{+2} and Mn^{+2} removal and increases filter working period.

RECOMMENDATIONS:

Using PP: PP is considered a very good selection for the process of iron and manganese removal. The experiments show that using dosages near to half of the required theoretical dose at normal pH and retention time of 20 minutes can remove iron completely and 90% of manganese. However, the dose should be determined very accurately to prevent water coloring.

Using sedimentation: sedimentation is required when iron and manganese

combined concentrations exceed 5.0 mg/l to allow longer filtration periods.

Using Alum: using alum with sedimentation alone cannot remove high ratio of manganese. However, using PP with alum affect Fe^{+2} removal. Using alum while raising pH to 10 leads to remove high ratio of both Fe^{+2}/Mn^{+2} concentrations.

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